

**Organics in the Northeastern Pacific and their impacts  
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# Organics in the Northeastern Pacific and their impacts on aerosol hygroscopicity in the subsaturated and supersaturated regimes

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## Abstract

Aerosol samples were collected by aircraft during the summer of 2004 in the North-eastern Pacific and compared to measurements of aerosol hygroscopicity. Chemical speciation analyses of the samples revealed that a significant portion of the marine aerosols was organic, and on average 8% of the total aerosol mass was insoluble organic material, tentatively attributed to natural marine emissions. Two chemical models were explored in an attempt to reproduce observed aerosol growth, both in the sub-saturated and supersaturated regimes. Results suggest that at subsaturated relative humidities, the nonideal behavior of water activity with respect to aerosol chemistry is an important effect on aerosol growth. At supersaturations, the underprediction of critical supersaturations by all models suggests the hypothesis that formation of a complete monolayer by the insoluble organics may inhibit the activation of aerosols to form cloud droplets.

## 1 Introduction

The global flux and burden of sea salt is comparable to anthropogenic sulfates and mineral dust, but the surface source of the marine aerosols and the washout removal mechanism limit their influence primarily to marine boundary layers. Global fluxes of sea salt are estimated to be  $\sim 1900 \text{ Tg yr}^{-1}$  in the Southern Hemisphere and  $\sim 1440 \text{ Tg yr}^{-1}$  in the Northern Hemisphere, with the difference between the two hemispheres due primarily to the greater land mass in the Northern Hemisphere. High trade winds in the northern midlatitudes generate a significant amount of marine aerosols, with yearly concentrations averaging as high as  $45 \text{ mg m}^{-2}$  within the marine boundary layer (Houghton et al., 2001). Thus in the remote marine atmospheres and some coastal areas, especially with high winds, the ocean constitutes an important source of aerosols. Observations have shown marine aerosols to be the prevailing source of light scattering and cloud condensation nuclei (CCN) in the remote ocean (O'Doud et al.,

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1997; Murphy et al., 1998; Quinn et al., 1998). Sea salt aerosols are an important part of many problems in atmospheric chemistry, atmospheric radiation, cloud physics and climate (Lewis and Schwartz, 2004).

Organics play an critical role in marine aerosol composition, with estimates that between 10–50% of the aerosol composition is organic (Maßling et al., 2003; Putaud et al., 2000). One source of organics in marine aerosols is photosynthetic marine organisms, which have been estimated to produce 30–60×10<sup>15</sup> g of organic carbon every year (Duarte and Cebrian, 1996). The carboxyl group of the long-chained fatty acids, a major constituent of the organics found in sea water, allows them to behave as surfactants, molecules containing a polar head and a hydrocarbon tail, such that these organics interact with water in contrasting ways. This surfactant nature causes them to accumulate at the ocean’s surface and, as marine aerosols are formed through a bubble bursting mechanism at the ocean’s surface, enrich the aerosol organic composition. Estimates of up to 200 Tg C yr<sup>-1</sup> of hydrocarbons available for incorporation in marine aerosol have been proposed (Ellison et al., 1999).

Other potential sources of organics in the marine aerosol include secondary organic aerosols produced through the oxidation of marine organics (Ellison et al., 1999) or transported pollutants (Kawamura and Sakaguchi, 1999). The oxidation process may take place in the gas-phase (Pun et al., 2000; Griffin et al., 2002), at the particle surface (Eliaison et al., 2003; Maria et al., 2004) or in the aqueous phase (Ervens et al., 2003; Warnick, 2003).

These internally mixed organics within the marine aerosols can potentially significantly impact the physical properties of marine aerosols and thus their global impact. Haywood et al. (1999) estimated a global clear-sky radiative impact of sea-salt due to scatter as -1.0 to -3.5 W m<sup>-2</sup>, and about 1.5 times greater over the oceans. Aerosol hygroscopicity, that is the change in diameter experienced by the aerosol due to water uptake associated with increasing relative humidity, plays a significant role in aerosol scatter, with chemical apportionment studies identifying condensed water as the most important contributor to ambient column optical depth off the Eastern coast of the

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United States (Hegg et al., 1997). The link between chemical composition and hygroscopicity is well established (e.g. Maßling et al., 2003; Winkler and Kaminski, 1992), but the role of organics in aerosol growth and activation is not well known, and indeed even their impact on the sign of globally averaged cloud radiative forcing is under debate.

## 2 Experiment

### 2.1 Data collection

The second Cloud and Aerosol Research in the Marine Atmosphere (CARMA II) campaign took place June and July of 2004, off the coast of Monterey, CA, as shown in Fig. 1. Several predominate features determine the climatology observed at the CARMA II campaign field site. Large-scale subtropical subsidence combined with cool oceanic surface temperatures create a well-defined marine boundary layer with low temperatures and high relative humidities within this mixed layer, resulting in a maximum in the stratocumulus cloud deck observed in the summer months. The northerly shift of the oceanic high pressure system in the summer months causes a shift to north-westerly winds, creating seasonal coastal ocean upwelling which increased the nutrient content of the surface waters near the California coast (Nuss, 1996). Due to the upwelling, the surface waters contain an increased nutrient content in those months which support a large phytoplankton community within the coastal waters, with dissolved organic carbon levels reaching as high as  $123 \mu\text{M}$  near the surface. The composition of the organics have not been completely characterized, and the constituents vary greatly, but they include amino acids (18–90% of organic carbon), wax esters (<1–20%), lipids (5–20%), tricycloglycerols (<1–30%), and fatty acids (<1–11%) (Coale et al., 1996).

The Center for Interdisciplinary Remotely-Piloted Aircraft Studies (CIRPAS) Twin Otter aircraft was the primary research platform used for data collection and housed a number of instruments to quantify meteorological conditions, cloud properties, gas phase concentrations and aerosol physical and chemical properties. The Twin Otter

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was equipped with an air intake probe used to sample a number of aerosol properties. This intake had a  $D_{50\%}$  transmission at approximately  $8\ \mu\text{m}$  (Hegg et al., 2005). The instruments employed in conjunction with the air intake probe are listed in Table 1. The Twin Otter instrumentation also includes a passive cavity aerosol spectrometer probe (PCASP): an external probe used to measure aerosol size-number concentrations (included in Table 1). Aerosol samples were collected on 47 mm Teflo<sup>TM</sup> filters and quartz filters using a filter sampler connected to the Twin Otter air intake probe.

## 2.2 Laboratory analysis

The laboratory analyses performed on the Teflo<sup>TM</sup> filters have been described in detail in previous papers (e.g. Gao et al., 2003), and thus will only be briefly described here. The aerosols collected on the Teflo<sup>TM</sup> filters were dissolved in a 10 ml aliquot of HPLC grade water, and underwent vigorous agitation to ensure even the moderately soluble species such as oxidized metals entered solution. The instruments used and the species identified, including the reproducibility, are listed in Table 2. The results of the Teflo<sup>TM</sup> filter analysis are shown in Fig. 2.

In addition to these techniques, thermal Evolved Gas Analysis (EGA) was performed on pre-fired quartz filters collected during the CARMA II campaign to determine the total organic content of the aerosols. Since the insoluble organic portion of the aerosols is of particular note, each quartz filter was cut in half after sampling and one half was thrice rinsed in HPLC grade water and dried using a standard vacuum aspiration technique to remove the soluble organic portion prior to EGA analysis. Each filter was associated with a backup filter on the second stage to account for positive carbon artifacts. EGA analysis was performed on a portion of each filter half with the temperature increasing at a constant rate from  $37^\circ\text{C}$  to  $737^\circ\text{C}$  in an oxidizing atmosphere in which the organics are converted to carbon dioxide with a catalyst present to ensure complete conversion. In-lab analysis found the reproducibility of this method to be within 16%.

In order to convert the organic carbon to soluble and insoluble particulate organic

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matter (SPOM and IPOM), representative molecular forms of SPOM and IPOM must be chosen to convert the organic carbon mass to total organic mass. Measurements made in this campaign and previous campaigns have suggested oxalic acid composes a significant portion of the resolved soluble organic mass in the marine atmosphere, and thus was chosen to represent soluble organic carbon (Kawamura and Sakaguchi, 1999; Crahan et al., 2004b). Similarly, palmitic acid was chosen to represent the insoluble portion of the organic fraction (Tervahattu et al., 2002). The results can be seen in Fig. 3.

On the occasions that the quartz filters were exposed simultaneously with one or two Teflo™ filters, the organic analyses results were added to the Teflo™ filter chemical speciation. To convert the resolved chemical speciation to total aerosol mass, total sea salt mass, in  $\mu\text{g m}^{-3}$ , was calculated using the equation:

$$\text{Sea Salt} = \text{Cl} + 1.47\text{Na} \quad (1)$$

such that chloride loss is accounted for, and the correction factor prior to the sodium term accounts for minor constituents in sea water such as calcium, potassium, magnesium, sulfate, and carbonate (Bates et al., 2005). Similarly, the trace metals were assumed to occur in their oxidized form and the total inorganic oxidized material (IOM) mass, in  $\mu\text{g m}^{-3}$  was calculated using:

$$\text{IOM} = 2.2\text{Al} + 2.49\text{Si} + 1.63\text{Ca} + 2.42\text{Fe} \quad (2)$$

This equation is based on the most common oxidized form of these metals found in soil and fly ash ( $\text{Al}_2\text{O}_3$ ,  $\text{SiO}_2$ ,  $\text{CaO}$ ,  $\text{K}_2\text{O}$ ,  $\text{FeO}$ ,  $\text{Fe}_2\text{O}_3$ ) (Seinfeld, 1986). Non-sea salt sulfate was calculated by subtracting the sea salt sulfate, 25.2% of the sodium mass, from the total sulfate mass. In order to find a lower bound for the total aerosol mass fraction composed of marine organics, complete neutralization of sulfate and nitrate by ammonium was assumed. Partial neutralization (i.e. the formation of ammonium bisulfate rather than ammonium sulfate) changes the total organic fraction by less than 2% on average. Assuming chemical closure, the results of the combined Teflo™ and quartz filters are shown in Fig. 4.

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### 3 Results and discussion

#### 3.1 Chemical speciation

There are several interesting points to note in the chemical speciation results. The total organic concentration of the marine aerosols averages to  $21\% \pm 14\%$  by mass, in agreement with previous estimates made globally (Maßling et al., 2003; Putaud et al., 2000). A significant portion of the total organic concentration is IPOM, constituting  $39\% \pm 24\%$  of the total POM. Another result of interest is that only a modest and highly variable  $12\% \pm 19\%$  of the total POM mass was speciated by the methods described above.

A correlation matrix was created to study the relationship between the various chemical species identified. The results are shown in Table 3. As can be seen, a strong linkage was found between the sulfate and nitrate, suggesting they are both due to transport of anthropogenic pollutants. Other strong correlations include one between the SPOM and the biomass burning tracers, suggesting biomass burning was a major source of soluble carbon during the campaign. There is no strong linkage between the dicarboxylic acids and any other chemical components, including most surprisingly the SPOM. However, this may be explainable as the dicarboxylic acids compose less than 20% of the SPOM. Furthermore, dicarboxylic acids within aerosols may have an additional, unique formation mechanism in cloud droplets (Warnick, 2003; Crahan et al., 2004a), in addition to the gaseous formation pathway, further complicating direct correlations. The lack of correlation between the IPOM and any other chemical marker is notable as well. This may be due to IPOM resulting from a complex mixture of both anthropogenic pollutants and natural marine sources. However, a multilinear regression performed on IPOM using sea salt and ammonium nitrate or sea salt and ammonium sulfate for marine and anthropogenic pollution yielded a poor correlation, with a correlation coefficient less than 0.3 in both cases. Back-trajectory analyses show that the sampled air parcels were over the ocean for two days or greater, minimizing the possibility that the IPOM had anthropogenic sources as the current estimated lifetime of

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IPOM is less than two days (Kanakidou et al., 2005). While true that a strong source such as a large forest fire might transport continental IPOM to the CARMA sampling area, IPOM, as shown in Table 3, is only slightly correlated with levoglucosan (0.229), so biomass burning is not the dominant source of IPOM. Another likely explanation is that, while the IPOM has a marine source, it varies both temporally and geographically compared to sea salt. Sea salt is well mixed in the ocean surface layer, and its emission from the ocean to form aerosols depends primarily on the wind and the resulting foam and spray. However, marine phytoplankton blooms primary production vary daily due to the nutrient content in the surface waters, and the strength of the blooms decrease with increasing distance from the coastal upwellings. Figure 5 shows both the chlorophyll content of the ocean surface waters, an indicator of the presence of marine phytoplankton, for the 28 June through the 2 July 2004 and the area over which the quartz samples were taken during that period, demonstrating the irregularity of the phytoplankton bloom strength over the sample area. Thus, sea salt is not a good candidate for a tracer of direct emissions of decomposition products of marine phytoplankton.

### 3.2 Impact of organics at subsaturated relative humidities

Organics potentially can impact aerosol growth thermodynamically in three manners: through the formation of an insoluble core in the hydrated aerosol, through impacting the water activity of the hydrated aerosol solution, or through lowering the surface tension of the water when behaving as a surfactant and gathering at the air-water interface of the hydrated aerosol.

Two methods are used in this paper to describe aerosol hygroscopic behavior, corresponding to the two instruments used to quantify aerosol growth. The UWPH measured the scatter of the sampled aerosols at a single wavelength ( $\lambda=540$  nm) at wet and dry relative humidities. The increase in scattering at the wet *RH* is related to the increase in aerosol diameter due to the uptake of water, as described by the Mie equation. The aerosol growth is then fit to the equation first described by Kasten (1969) (see also:

Gassó et al., 2000):

$$\sigma_{sp} = k \left( 1 - \frac{RH}{100} \right)^{-\gamma} \quad (3)$$

where  $\sigma_{sp}$  is the particle scattering coefficient,  $RH$  is the ambient relative humidity, and  $k$  and  $\gamma$  are determined by fitting the equation to known data. The  $k$  depicts the  $\sigma_{sp}$  at dry  $RH$ s and  $\gamma$  describes the hygroscopic growth of the aerosol with higher  $\gamma$  values associated with more hygroscopic aerosols (i.e. those whose light scattering is more dependent on  $RH$ ). Assuming the hydrated aerosols at the two measured  $RH$ s remain on the upper branch of the hysteresis curve, the measurement of scatter at two  $RH$ s by the UWPH eliminates the need to determine the  $k$  parameter prior to the determination of the  $\gamma$  parameter. This  $\gamma$  parameter is most relevant to modeling the direct global radiative impacts of aerosols.

The  $\gamma$  values and (mean and one standard deviation) were determined for the data set containing complete Teflo<sup>TM</sup> and quartz filter results, and the results are plotted on Fig. 4. Also included in Fig. 4 are the mole fractions of the chemical constituents of the aerosols. Correlations were performed between the chemical mole fractions and the  $\gamma$  values, and the strongest relationship, with a linear correlation coefficient of 0.85, was found between the sea salt mole fraction and the  $\gamma$  values, demonstrating the strong linkage between aerosol chemical components and aerosol hygroscopicity. Additionally, a strong negative correlation was observed between the  $\gamma$  values and the IPOM mole fraction, with a linear correlation coefficient of  $-0.78$ . Thus, while the IPOM composes on average only 7% of the total aerosol composition, it plays a strong role in aerosol hygroscopicity, and should be included in aerosol hygroscopicity models. Surprisingly, a strong negative correlation ( $-0.70$ ) was found between the  $\gamma$  values and the ammonium sulfate fraction. This can be explained in that ammonium sulfate, while a hygroscopic compound, is less hygroscopic than sea salt and thus may have a relatively depressing effect on the total marine aerosol hygroscopicity.

The H-TDMA also measured aerosol hygroscopicity by determining the diameter growth of aerosols at an  $RH=85\%$  from  $RH<20\%$  from four different dry diameters

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( $d=50$  nm, 100 nm, 150 nm, 200 nm). The diameter growth factor (DGF) is then calculated as the ratio of the wet diameter ( $d_{\text{wet}}$ ) to the dry diameter ( $d_{\text{dry}}$ ):

$$\text{DGF} = \frac{d_{\text{wet}}}{d_{\text{dry}}} \quad (4)$$

To further explore the relationship between aerosol chemical composition and aerosol hygroscopicity, three cases were studied using both H-TDMA results and quartz filter analyses results. The H-TDMA time series for the two days, and the periods for the case studies (labeled C1, C2, and C3) are shown in Fig. 6. Attempting to reproduce the observed DGF for these three case studies, two models were chosen: the Köhler model using five different initial assumptions, and the Svenningsson model (Svenningsson et al., 2005) using two sets of initial assumptions. As the H-TDMA measured the DGF at four dry aerosol sizes, a total of twelve points will be compared to model calculations.

The Köhler model is the traditional model used to explain aerosol hygroscopic growth, where the water uptake is dependent upon two parameters, the Kelvin parameter ( $a$ ) and the solute parameter ( $b$ ). Equation (5) describes the relationship between these two effects, where  $R_p$  is the radius of the hydrated aerosol at a given saturation ( $S$ ), and  $r_p$  is the radius of the insoluble core of the aerosol.

$$S = \frac{a}{R_p} - \frac{b}{(R_p^3 - r_p^3)} \quad (5)$$

$$a = \frac{2\sigma}{\rho_l R_v T} \quad (6)$$

$$b = \frac{3im_l M_s}{4\pi\rho_l m_s} \quad (7)$$

The Kelvin effect, describing the change in the surface tension due to the curvature of the droplet, is calculated using the standard surface tension of the liquid in a plane surface ( $\sigma$ ), the density of water ( $\rho_l$ ), the surrounding temperature ( $T$ ) and the gas

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constant for dry air ( $R_v$ ). The solute effect is the effect of the species in solution on the hydrated aerosol. It is dependent upon the mass of the species in solution ( $M_s$ ) and its molecular weight ( $m_s$ ), as well as the number of ions it forms when dissolved in solution ( $i$ ). A major assumption of the solute effect of the Köhler equation is that the solution is dilute enough that the water activity is ideal and that the impact of each additional species is additive and linear. There are several potential ways that SPOM and IPOM can impact the Köhler terms. While unlikely, it is possible that everything will enter solution. The IPOM may form an insoluble core in the hydrated aerosol, and the SPOM may either enter solution or form an insoluble core. It is more likely that the SPOM will be in solution, but several studies have suggested that the partitioning of SPOM between the aqueous phase and the aerosol core may have a significant impact on aerosol formation, and by allocating all of the SPOM to the core, the results should bound one extreme of this effect (e.g. [Bilde and Svenningsson, 2004](#)). A final possibility is that the IPOM may act as surfactants, gathering at the air-water interface and reducing the surface tension.

To explore these possibilities, the Köhler model was run several times. A well-mixed aerosol was assumed, and the chemical constituents were simplified following a convention suggested by [Svenningsson et al. \(2005\)](#), such that the sea salt was assumed to be in the form of NaCl, levoglucosan ( $C_6H_{10}O_5$ ) was used as a tracer for biomass burning, succinic acid ( $C_3H_6O_4$ ) was used as a tracer of SPOM and fulvic acid ( $C_{33}H_{33}O_{19}$ ) was used as a tracer for IPOM. The IOM fraction of the aerosol composition for the three case studies ranged from 0% to 4%, and when present, assumed to constitute an insoluble core within the aerosol. The van't Hoff constants assumed for the inorganic species were 2, 2, and 2.2 for ammonium nitrate, sodium chloride and ammonium sulfate, respectively, and one for all organics, as determined by [Svenningsson et al. \(2005\)](#) through laboratory measurements. In the case where the IPOM behaved as a surfactant, the surface tension was set to remain constant at 44 dynes, the lowest surface tension observed for fulvic acid in laboratory measurements ([Svenningsson et al., 2005](#)).

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The Svenningsson model is based upon laboratory measurements of four idealized aerosol mixtures, sea salt, biomass burning, pollution, and organic. The sea salt mixture assumed is similar to those found in the three case studies, as shown in Table 4. The Svenningsson model is a set of equations that were fitted to the results of the laboratory measurements and used to describe the change in water activity with changing total solution molality, and the change in surface tension to the change in carbon solution molality. This allowed both for the decreasing impact of surfactants on the solution as the aerosol diameter grew, but also allowed for the nonideal behavior of the water activity since the change observed in water activity with aerosol growth was nonlinear. The Svenningsson model results were calculated twice, once using the full set of equations suggested by the model, and once keeping the surface tension constant at 76 dynes, that of pure water.

The best and worst agreement for the Köhler model runs are shown in Fig. 7, as well as the results of the Svenningsson model runs. The error analyses for the Köhler curves are incomplete, and only based on uncertainty in the chemical analysis reproducibility. The errors in this application of the Svenningsson model are calculated based upon errors suggested by Svenningsson et al. (2005). The results for all the model runs are summarized in Table 5.

The largest H-TDMA errors are caused by uncertainties in the internal  $RH$ . The  $RH$ s upstream and downstream of the second DMA were closely monitored and controlled, but the inner cylinder is slow to respond, which results in an elevated (but unknown)  $RH$ . The magnitude of this error decreases during flight and is probably negligible by the end of each flight. This error would be time-dependent and coupled to the cabin temperature of the aircraft, which is different for each flight, making even a guess at this uncertainty difficult. Future versions of the aircraft H-TDMA are attempting to address this issue. However, ultimately this error is likely to be small.

The reader should be aware that the aerosols sampled during the CARMA-II campaign were most likely not well-mixed aerosols, as assumed during the model initializations. However, hygroscopic growth factor distributions measured by the H-TDMA

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were monomodal, suggesting that at least particles of a given size had relatively uniform composition. Furthermore, there are significant advantages to the bulk analysis method used above as this method is simpler to translate to modeling efforts. The relatively small variations in the  $\gamma$  values observed for similar air masses during the CARMA II campaigns, as seen in Fig. 4, support the idea that variations between aerosols of different sizes are less important when sampling bulk aerosol properties.

The results show that the assumptions used in the Svenningsson model are clearly superior to those used in the Köhler model for the marine aerosol case studies. This suggests that treating water activity as ideal is not a good assumption at subsaturations where the solution is concentrated. Another interesting result is that changing the surface tension of the solution does not significantly change the agreement between the calculated and the observed DGF within the Svenningsson model. The nonideal water activity is a much more significant effect, warranting further research. Organics likely play a large role in the poor agreement between the models and observations, as SPOM and IPOM can significantly depress the water activity at high  $RH$ s (Ming and Russell, 2001).

### 3.3 Impact of organics at supersaturated relative humidities

Organics play a potentially major role in aerosol activation to form cloud droplets. The depression in water activity due to the organics will increase the critical supersaturation necessary to activate the aerosols. Additionally, two competing effects have been suggested for the role of IPOM in aerosol activation. Facchini et al. (2000), among others, theorized that the reduction in surface tension caused by the surfactants decreases the critical supersaturation for the aerosols and increase cloud droplet number concentrations by as much as 20%. However, Ellison et al. (1999) suggested that the IPOM may create an inverted micelle, with the polar heads of long-chained surfactants pointing inward towards the aerosol solution and the nonpolar tails pointing outwards. If sufficient organics are present, this tightly packed monolayer may impede evaporation or condensation of water, with delays of up to 30 s for aerosol activation observed in

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the laboratory for an aerosol coated in cetyl alcohol within a CCN static diffusion chamber (Bigg, 1986). Similar behavior has been observed for internally mixed ammonium sulfate/organic aerosols by Shantz et al. (2003) and Abbatt et al. (2005). However, once growth begins and the surfactant shell is broken, water uptake proceeds without measurable inhibition (Podzimek and Saad, 1975). Chuang et al. (1997) calculated the potential impact of delays in activations on clouds, concluding that neglecting this impact in calculations could overestimate the cloud droplet concentration in clouds by as much as 69%.

The Köhler equations and the Svenningsson model calculations can be extended to predict the critical supersaturations of the aerosols, and compared to observations made by the cloud static thermal diffusion chamber during filter collection times. To screen out instrument noise, the data was fit to the equation:

$$\text{CCN}(s) = cs^k \quad (8)$$

where CCN is the cloud droplet number concentration, and is a function of supersaturation ( $s$ ). The parameters  $c$  and  $k$  are constants determined by the aerosol type and size distribution. The cloud chamber data at four supersaturations (0.18, 0.30, 0.34, 0.50) were fit to the equation and only the fits at the 95% confidence level were retained. By applying this filter, one of the case studies was eliminated.

For the remaining two case studies, a 50% activation efficiency was assumed to indicate activation within the cloud chamber. To determine if the critical supersaturation was reached, the PCASP was used to calculate the aerosol number concentration at the dry diameter. The 50 nm critical supersaturations were not calculated because the PCASP lower boundary was 100 nm. As the internal  $RH$  of the PCASP was not completely dry, a uniform DGF of 1.1 (Ming and Russell, 2001) was assumed at  $RH=30\%$  and the aerosol distribution was adjusted downward accordingly. The aerosol number concentrations above the lower boundary diameters of 100 nm, 150 nm and 200 nm were estimated from the PCASP size distribution for C1 and C3, as shown in Fig. 8. When the ratio of the number of activated cloud droplets in the cloud chamber to

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PCASP number concentration above the lower boundary diameter was greater than 0.5, then it was assumed that critical supersaturation was reached for aerosols at and above that diameter.

The results of the critical supersaturation calculations, compared to the observations are shown in Fig. 9. The bottom error bars for the observed critical supersaturations extend to the supersaturation in the prior chamber as it cannot be certain where within the range of supersaturations between the two chambers the aerosol population activated. Within these two case studies, the Svenningsson model comes closest to predicting the observed supersaturation but still significantly underpredicts observations. There are two explanations for this discrepancy. One possibility is that the supersaturation in the cloud chamber for a given temperature setting shifted from the values obtained during calibration. In this regard, it should be noted that the chamber was not calibrated during the CARMA II campaigns, but rather before shipment to the CARMA II site, and the shipment may have induced such a shift. Another potential explanation is the formation of the inverted micelle. A delay of 30 s would be sufficient to prevent activation of the aerosols in the cloud chamber, as the residence time in the cloud chamber is only 20 s. It is possible that a kinetic effect was not observed in the H-TDMA measurements because kinetic effects would have a lesser impact on the hygroscopicity measurements because the growth timescales of  $\sim 0.2 \mu\text{m}$  (H-TDMA) and  $\sim 2 \mu\text{m}$  (CCN chamber) particles are very different.

To explore the possibility of the formation of a monolayer, the PCASP surface area was calculated for the aerosol population for the combined filter results. To do this, first the reliability of the PCASP to accurately measure the aerosol population must be established, so a mass closure study was performed. Aerosol mass was estimated from PCASP volume concentrations using a dry density of  $1.9 \text{ g cm}^{-3}$ , an aerosol refractive index of 1.41, and a DGF=1.1 at  $RH=40\%$ . The density was chosen based upon the mass-weighted average density of the components listed in Table 4. An error in aerosol density of  $\pm 0.2 \text{ g cm}^{-3}$ , and of a refractive index  $\pm 0.16$ , as well as the standard error of the PCASP measurements over the measurement period were propagated through

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the equation using Taylor series expansion and included in the error bars. The PCASP has an upper bound of 3.6  $\mu\text{m}$  and reliable measurements of the aerosol size-number concentration do not exist beyond that diameter for the CARMA II campaign. It was found that the PCASP mass concentrations agreed with the mass of the speciated aerosol within the error bounds for all cases except for one, as shown in Fig. 10. Using the PCASP surface area measurements, with the upper bound of 3.6  $\mu\text{m}$ , is likely sufficient for determining if there is a monolayer as previous studies have found that marine IPOM is concentrated in the submicron range (O'Doud et al., 2004; Cini and Loglio, 1997), due to the formation mechanism of sea salt aerosols (Oppo et al., 1999).

To calculate a surface area coverage of the IPOM, palmitic acid was used as a representative compound with a surface area of 20  $\text{\AA}^2$  for a single molecule, since a tightly packed monolayer is necessary to impede evaporation and condensation of gases (Seidl, 2000). The results were compared to the PCASP surface area measurements, and shown in Fig. 11. A surface area fraction at or above one is a complete monolayer. As can be seen in Fig. 11, there is excess IPOM for the formation of the monolayer at  $RHs$  experienced in the PCASP. In order to account for aerosol growth at ambient  $RHs$ , the PCASP aerosol volume of the three case studies were hydrated using the average  $DGF=1.5$  measured by the H-TDMA and the surface area was recalculated. The results are included in Fig. 11. Even at ambient  $RHs$ , the marine aerosols observed during the CARMA II campaigns may be covered in a tightly packed monolayer of IPOM. Thus, this is a possible explanation for the discrepancy observed between modeled critical supersaturations and those observed onboard the Twin Otter for the two case studies.

## 4 Conclusions

The linkage between aerosol chemistry and aerosol hygroscopicity and CCN activity is strong. Analyses of the  $\gamma$  values observed during the CARMA II campaigns showed the strong effect that sea salt and IPOM has on aerosol hygroscopicity. However, the

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traditional Köhler model used to describe the relationship between the two does not account for the impact that organics have on aerosol hygroscopicity. Organics can suppress the water activity of a solution in a nonlinear fashion, decreasing the equilibrium  $RH$  above the concentrated solution of a hydrated aerosol. Three case studies explored using chemical and hygroscopic data collected during CARMA II demonstrated that the best agreement between predicted and observed aerosol growth came when nonidealities were taken into account in the model. This effect was found to be much stronger than potential changes in the surface tension caused by surfactants, suggesting that more research should be focused on describing nonideal water activity when predicting aerosol growth using chemical constituents.

Significant discrepancies were identified between the thermodynamically predicted and observed critical supersaturations for several aerosol case studies. While calibration issues cannot be entirely eliminated as explanations, sufficient IPOM was found in the aerosol chemical composition to form a complete monolayer that may impede activation of the aerosols to form cloud droplets. As the amount of time a marine aerosol particle spends exposed to maximum supersaturation in-cloud is brief, the delay in cloud condensation nuclei activation could impact the number of droplets and the resulting cloud albedo.

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**Table 1.** The onboard instrumentation aboard the CIRPAS Twin Otter during RED Campaign. Refractive index is symbolized by r.i.

| Instrument   | Model                                     | Values Measured                    | Parameters  |
|--|---|------------------------------------|---|
| UW passive humidigraph (UWPH)                              | Model 903, Radiance Research, Seattle, WA | Light scattering ( $\sigma_{sp}$ ) | $\lambda=540$ nm,<br>$RH \approx 25\%, 50\%, 80\%$                        |
| TSI nephelometer   | Model 3563, TSI, St. Paul, MN             | Light scattering ( $\sigma_{sp}$ ) | $\lambda=450$ nm,<br>550 nm, 700 nm                                       |
| Cloud static thermal diffusion chamber                     | U. Wyoming                                | Critical Supersaturation           | $s=0.18, 0.30,$<br>0.34,0.50  |
| Hygroscopic Tandem Differential Mobility Analyzer (H-TDMA) | Aerosol Dynamics, Inc., Berkeley, CA      | Diameter growth factor             | $d_{dry}=50$ nm, 100 nm,<br>150 nm, 200 nm;<br>$RH(1) < 20\%, RH(2)=85\%$ |
| Passive Cavity Aerosol Spectrometer Probe 100x (PCASP)     | PMS/DMT Inc., Boulder, CO                 | Aerosol size/number distribution   | $d=0.100\text{--}3.17$ nm;<br>$RH \approx 30\%$<br>r.i.=1.41              |

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**Table 2.** Instruments used for the laboratory analyses of aerosols collected on the Teflo™ filters. The reproducibility of each species is in parentheses next to the species.

| Instrument  | Model  | Species identified   |
|---|--|--|
| Ion Chromatograph (IC)  | DX-500,<br>Dionex,<br>Sunnyvale, CA                      | Cl <sup>-1</sup> (5%); NO <sub>3</sub> <sup>-2</sup> (4%); SO <sub>4</sub> <sup>-3</sup> (3%);<br>formate (27%); glutarate (10%);<br>succinate (3%); malonate(14%);<br>oxalate (15%) |
| Inductively Coupled<br>Plasma Atomic<br>Emission Spectrometer (ICP-AES)   | Model ASU995,<br>Jarrell Ash,<br>Grand Junction, CO      | Mn(2%); Zn(2%); Ca(4%);<br>Mg(3%); Na(2%);<br>Fe(2%); K(1%); Si(3%); Al(1%)  |
| Electrospray ionization-ion<br>trap mass spectrometer (MS)<br>coupled with IC in<br>pulsed amperometric detector<br>mode (IC-PAD) | Esquire-LC, Bruker/<br>Hewlett-Packard,<br>Billerica, MA | levoglucosan(2%); glucose(11%)   |

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**Table 3.** Correlation matrix for the aerosol components speciated using the Teflo™ and quartz filter results. DA stands for dicarboxylic acids (e.g. oxalic acid, malonic acid, glutaric acid and succinic acid).

|   | (NH <sub>4</sub> ) <sub>3</sub> SO <sub>4</sub> | (NH <sub>4</sub> )NO <sub>3</sub> | Sea salt | IOM    | Biomass | DA     | Other SPOM | IPOM  |
|---|---|-----------------------------------|----------|--------|---------|--------|------------|-------|
| (NH <sub>4</sub> ) <sub>3</sub> SO <sub>4</sub> | 1.000   |                                   |          |        |         |        |            |       |
| (NH <sub>4</sub> )NO <sub>3</sub>               | 0.920   | 1.000                             |          |        |         |        |            |       |
| Sea salt  | 0.437   | 0.653                             | 1.000    |        |         |        |            |       |
| IOM   | -0.235  | -0.007                            | 0.113    | 1.000  |         |        |            |       |
| Biomass   | -0.018  | -0.079                            | 0.369    | -0.216 | 1.000   |        |            |       |
| DA  | 0.220   | 0.501                             | 0.575    | 0.482  | -0.055  | 1.000  |            |       |
| Other SPOM                                      | 0.014   | -0.089                            | 0.144    | -0.107 | 0.624   | -0.376 | 1.000      |       |
| IPOM  | 0.182   | 0.050                             | -0.085   | -0.096 | 0.229   | -0.025 | -0.026     | 1.000 |

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**Table 4.** A comparison of the Svenningsson experimental sea salt fractional composition to that observed in the three CARMA II case studies.

| Component        | Svenningsson | C1   | C2   | C3   |
|------------------|--------------|------|------|------|
| X(Levogluconan)  | 0            | 0    | 0.0  | 0.02 |
| X(Succinic Acid) | 0.10         | 0.07 | 0.13 | 0.17 |
| X(Fulvic Acid)   | 0.10         | 0.04 | 0.03 | 0.04 |
| X(Amm. Nitrate)  | 0            | 0.10 | 0.06 | 0.04 |
| X(Amm. Sulfate)  | 0.50         | 0.41 | 0.21 | 0.28 |
| X(Sea Salt)      | 0.30         | 0.38 | 0.53 | 0.45 |
| X(IOM)           | 0            | 0    | 0.04 | 0    |

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**Table 5.** Results of the Köhler and Svenningsson model runs compared to H-TDMA measurements.

| Model        | Surface Tension<br>( dynes)       | Water activity          | Core            | Average Absolute Error |
|--------------|-----------------------------------|-------------------------|-----------------|------------------------|
| Köhler       | 76                                | Ideal                   | IOM, if present | 35%                    |
| Köhler       | 76                                | Ideal                   | IPOM, IOM       | 36%                    |
| Köhler       | 76                                | Ideal                   | POM, IOM        | 34%                    |
| Köhler       | 44                                | Ideal                   | IOM, if present | 38%                    |
| Köhler       | 44                                | Ideal                   | SPOM, IOM       | 36%                    |
| Svenningsson | 44–76<br>(varies with C molality) | Varies<br>with molality | None            | 15%                    |
| Svenningsson | 76                                | Varies<br>with molality | None            | 13%                    |

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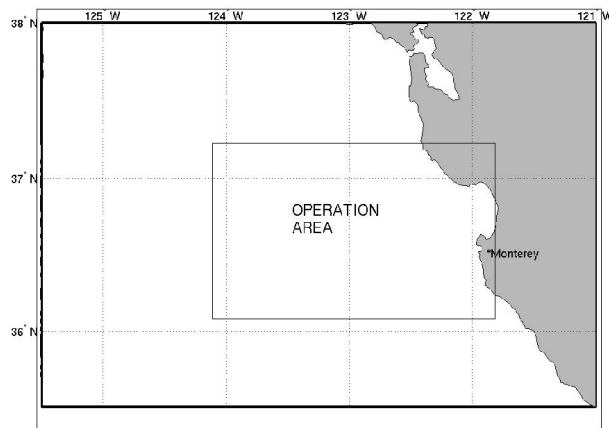
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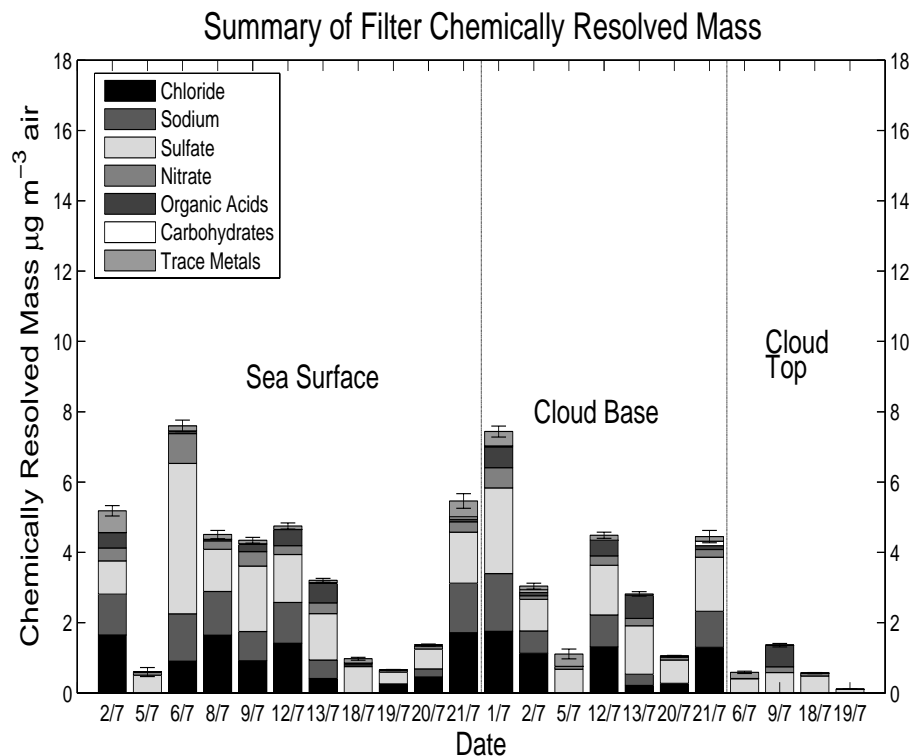


**Fig. 1.** A map of showing the operation area of the CARMA II field campaign off the coast of Monterey, California.

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**Fig. 2.** Results of chemical speciation analyses of the aerosols collected aboard the Twin Otter using Teflo<sup>TM</sup> filters. Results are divided both by the day which they were collected and the level collected within the marine boundary layer.

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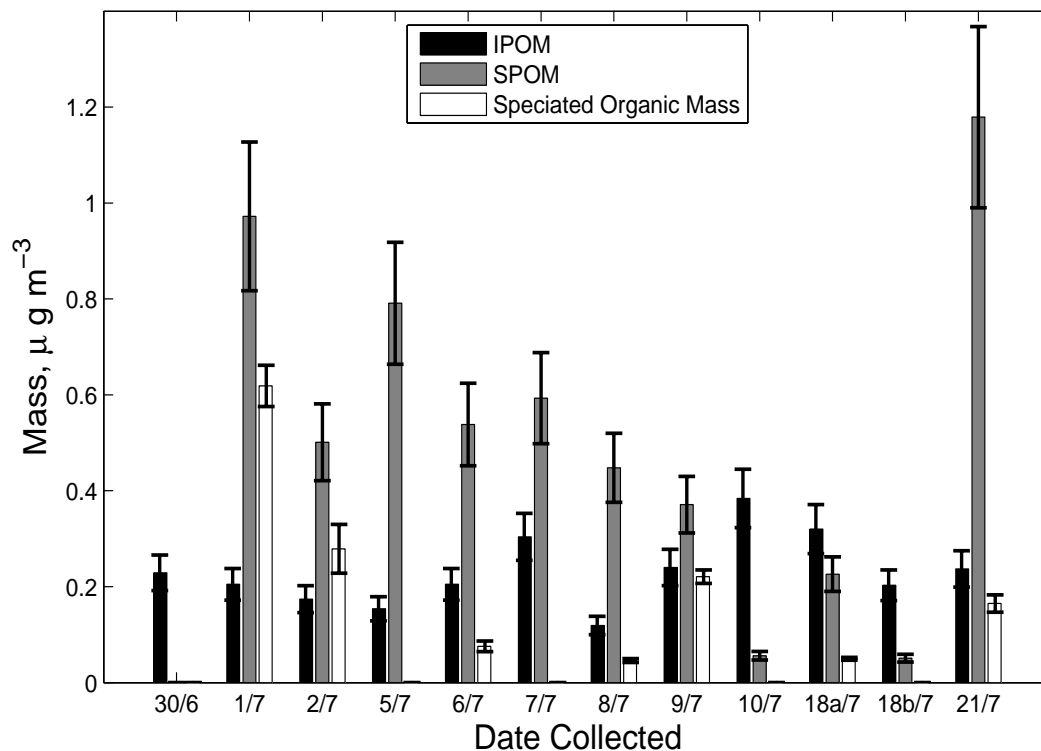
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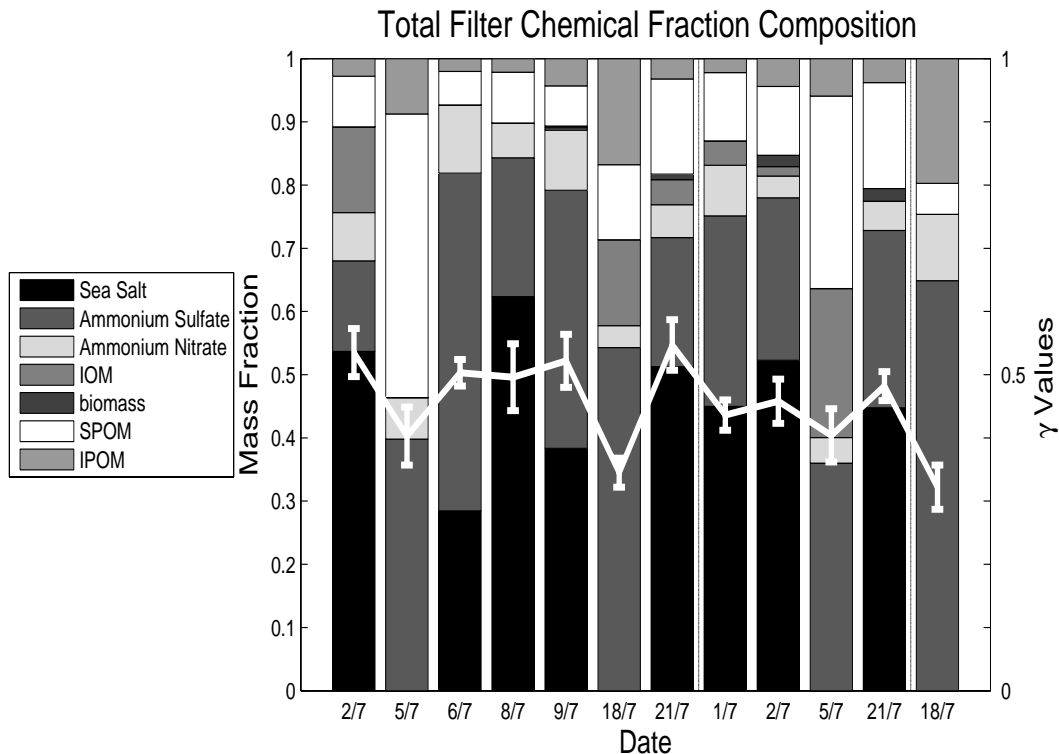


**Fig. 3.** The results of carbon EGA analysis of the aerosols collected aboard the Twin Otter using quartz filters. Oxalic acid and palmitic acid were used to convert the soluble and insoluble portion, respectively, of the organic material to POM. Also included are the speciated carbon results from the Teflo<sup>TM</sup> filters.

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**Fig. 4.** Chemical closure results for the combined Teflo<sup>TM</sup> and quartz filter results. Results are separated both by date and level collected. The first seven filters were collected at sea surface, the next four at cloud base, and the last one at cloud top. The  $\gamma$  values determined for the filters are overlaid on the chemical closure results.

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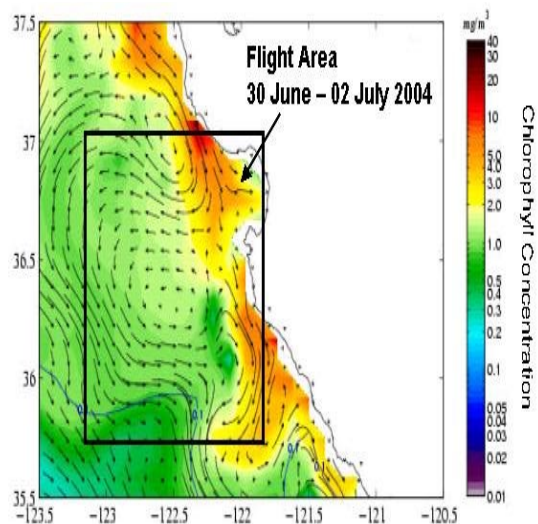
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**Fig. 5.** A model reproduction of satellite data of chlorophyll concentrations and sea surface currents over the sample areas from the 28 June to 2 July 2004 (Kindle, 2005) showing the geographic variability of chlorophyll, a marker for phytoplankton blooms. The black box indicates the bounds of the flight path during the quartz filter sample periods for samples collected on the 30 June and the 1 and 2 July.

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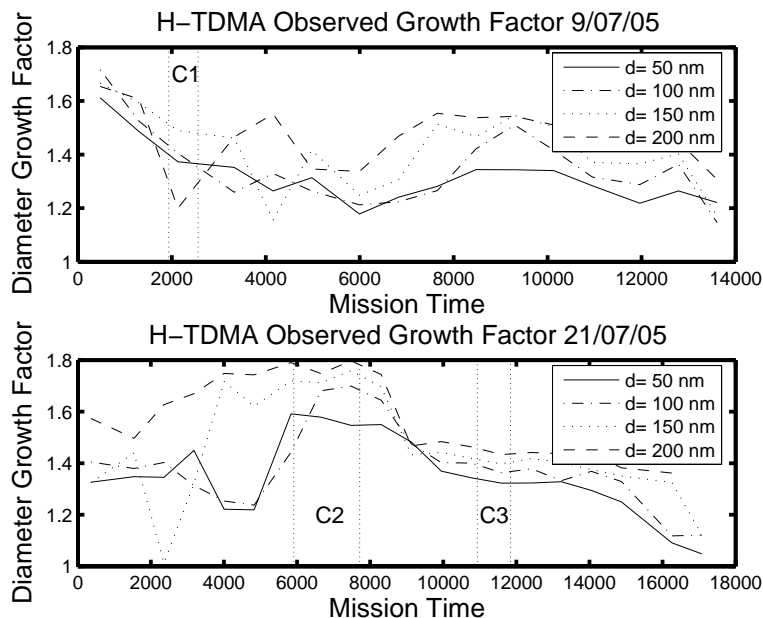
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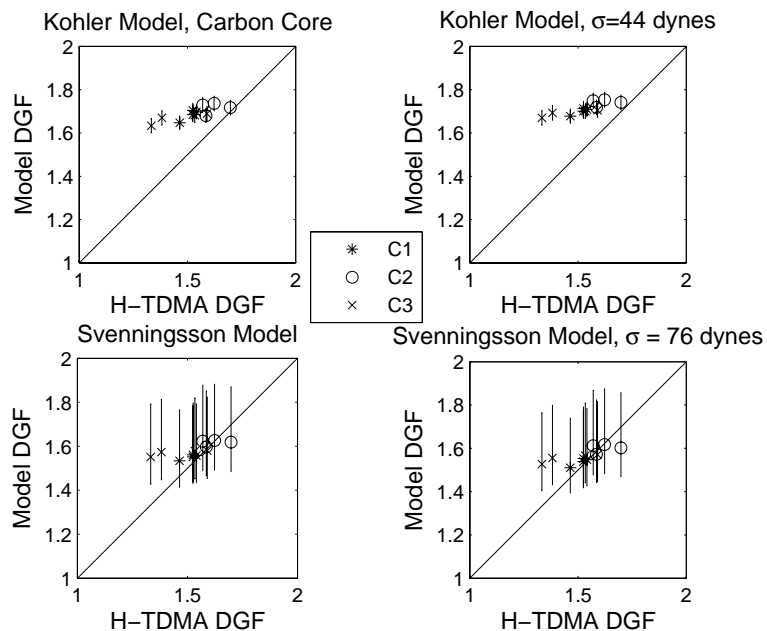


**Fig. 6.** The time series for the H-TDMA measured diameter growth factor for two mission flights. The time period of the three case studies (C1, C2, and C3) are indicated.

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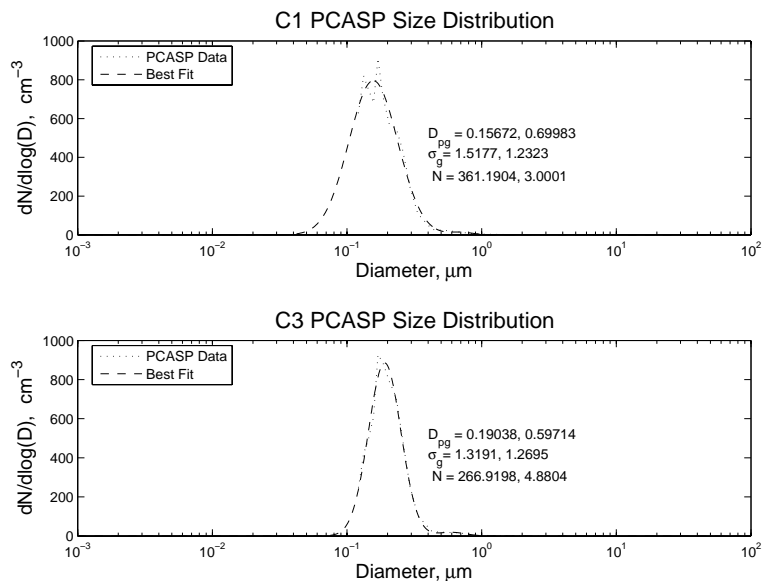


**Fig. 7.** The results of the Köhler model calculation and the Svenningsson model calculations compared to the H-TDMA measured DGF.

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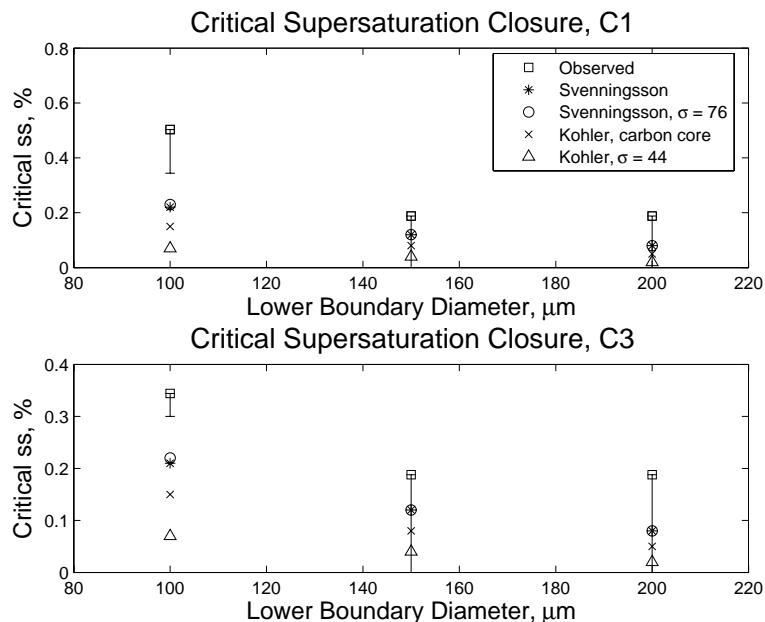


**Fig. 8.** The observed and fitted aerosol accumulation mode number concentration for Case 1 and Case 3.

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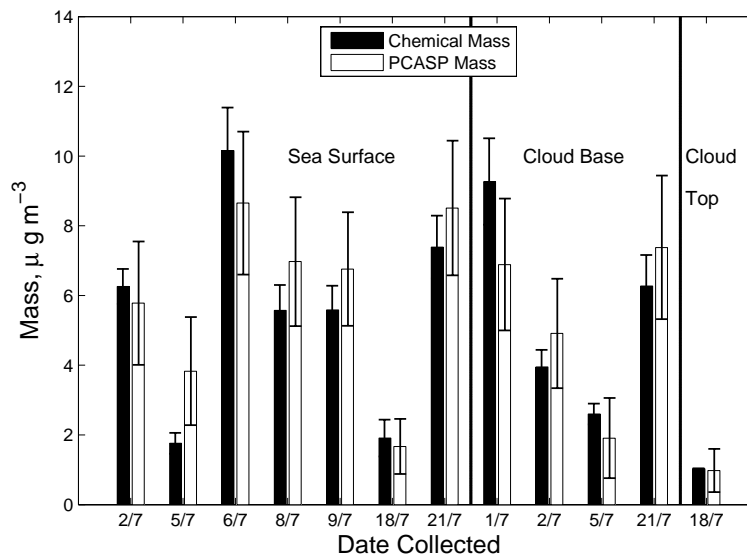


**Fig. 9.** The observed and calculated CCN critical supersaturations (ss) for C1 and C3, using the static cloud thermal diffusion chamber and the Köhler and Svenningsson model calculations.

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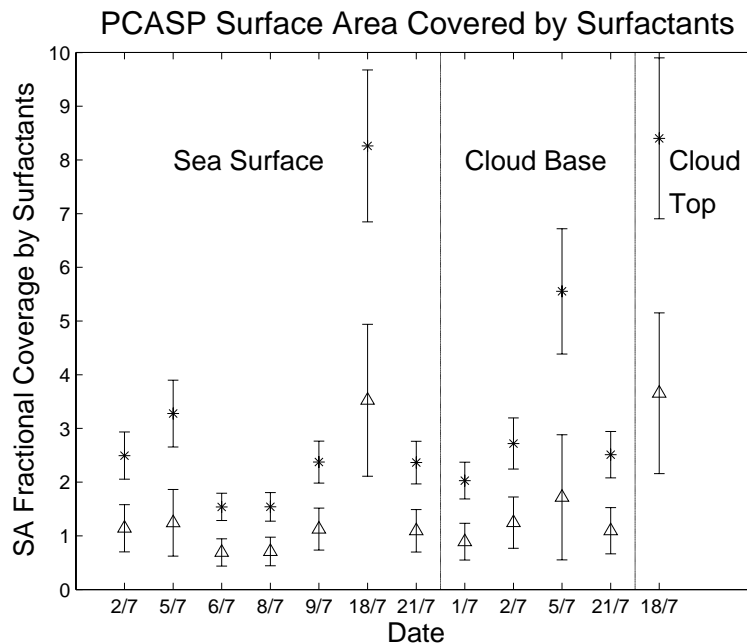


**Fig. 10.** A comparison of chemically resolved aerosol mass and mass observed by the PCASP.

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**Fig. 11.** The ratio of palmitic acid surface area to PCASP observed surface area for aerosols collected using quartz filters. The stars show the surface area coverage at a dry  $RH$  and the triangles show the surface area coverage at  $RH=85\%$ , assuming an average  $DGF=1.5$ . A ratio of one suggests a complete monolayer is present.

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